

Name: \_\_\_\_\_

Period: \_\_\_\_\_

**HW3:6 Guided Reading 2**  
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**Assigned: Fri., 10/20 and Mon., 10/23**  
**Due: Tues., 10/24 and Fri., 10/25**

*Start on page 119.*

1. Do covalent compounds have high or low melting points?

2. What is a nonpolar molecule?

3. What do you know about water if water is a polar molecule? (figure it out from the discussion about nonpolar molecules or look up water in the index)

4. Why do 2 Chlorines share one electron each?

5. How many electrons do Oxygen and nitrogen share?

*On p.120 it tells you what kind of elements electrons are more attracted to (a certain area of the periodic table).*

6. Would electrons be more attracted to Fluorine or Bromine?

7. Would electrons be more attracted to Nitrogen or Oxygen?

*Using the chart bottom of p.124*

8. What is the oxidation # for Copper (I)?

9. What is the oxidation # for Nickel (III)?

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NOT FROM THE READING (Though it would help).

10. How many electrons are gained or lost by Oxygen?

11. How many electrons are gained or lost by Sodium?

12. How many electrons would be gained or lost by 2 Sodium atoms?

13. How many electrons would be gained or lost by 2 Aluminum atoms?

*Reading again, p. 135*

14. What do we call the shape of DNA (ladder is wrong):

15. C always pairs with:

16. A always pairs with:

17. Give the opposite letters for in the opposite side of this DNA molecule.

