

Card 1

Chapter 19

Question:

The outer electrons in an atom are called:

Card 2

Chapter 19

Question:

A negative ion. Occurs when a non-metal gains electrons (more electrons than protons)

Card 3

Chapter 19

Question:

How many electrons does F^{1-} have?

Card 4

Chapter 19

Question:

Write ion notation for the element with 15 protons and 18 electrons

Card 5

Chapter 19

Question:

A bond between two non-metals when atoms share electrons. Requires the use of prefixes when naming.

Card 6

Chapter 19

Question:

Which element on the left side is actually a non-metal (the one exception)

Card 7

Chapter 19

Question:

A molecule composed of two or more elements

Card 8

Chapter 19

Question:

Write ion notation for the element with 12 protons and 10 electrons

Card 2

Chapter 19

Answer:

anion

Card 1

Chapter 19

Answer:

Valence electrons.

Card 4

Chapter 19

Answer:

P^{3+}

Card 3

Chapter 19

Answer:

10 electrons
(9 protons - 10 electrons = 1- charge)

Card 6

Chapter 19

Answer:

hydrogen

Card 5

Chapter 19

Answer:

covalent bond (ex. Carbon dioxide)

Card 8

Chapter 19

Answer:

Mg^{2+}

Card 7

Chapter 19

Answer:

compound
(H_2 is a molecule; H_2O is a compound)

Card 9

Chapter 19

Question:

A bond between a metal and a non-metal,
between a positive ion and a negative ion.
Electrons are actually given and taken.

Card 10

Chapter 19

Question:

The elements on the border between metals
and non-metals. They have properties of both.

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Chapter 19

Question:

A group of atoms with a net charge
(a group ion)

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Chapter 19

Question:

Write ion notation for the element with
7 protons and 10 electrons

Card 13

Chapter 19

Question:

An atom with a net electrical charge of zero
(number of protons = number of electrons)

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Chapter 19

Question:

Metal or non-metal: Fluorine? Calcium?

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Chapter 19

Question:

A compound with water
chemically bonded to it

Card 16

Chapter 19

Question:

What process do you use to
make ionic compounds?

Card 10

Chapter 19

Answer:

metalloids or semi-metals (ex. B, Ge, Si, As)

Card 9

Chapter 19

Answer:

ionic bond (ex. Sodium chloride)

Card 12

Chapter 19

Answer:

N^{3-}

Card 11

Chapter 19

Answer:

polyatomic ion

Card 14

Chapter 19

Answer:

Fluorine is a non-metal; calcium is a metal

Card 13

Chapter 19

Answer:

neutral atom

Card 16

Chapter 19

Answer:

Cross the number not the sign.
(The numbers are oxidation numbers.)

Card 15

Chapter 19

Answer:

hydrate

Card 17

Chapter 19

Question:

An atom with a positive or negative charge
(different number of protons than electrons)

Card 18

Chapter 19

Question:

Atoms are more stable when they have a
full set of valence electrons, usually eight.
This is known as what rule?

Card 19

Chapter 19

Question:

An element that exists in pairs.

Card 20

Chapter 19

Question:

When making dot diagrams
what do the dots represent?

Card 21

Chapter 19

Question:

Elements on the left side of the periodic table

Card 22

Chapter 19

Question:

What charge will Oxygen have if
it gains two electrons?

Card 23

Chapter 19

Question:

What property of an element most determines
its chemical characteristics?

Card 24

Chapter 19

Question:

In dot diagrams do you
put the dots in a circle?

Card 18

Chapter 19

Answer:

octet rule (if | 8 | full)

Card 17

Chapter 19

Answer:

ion

Card 20

Chapter 19

Answer:

Valence electrons.

Card 19

Chapter 19

Answer:

diatomic molecule (O_2 , N_2 , H_2)

Card 22

Chapter 19

Answer:

2- (gainers of electrons become negative)

Card 21

Chapter 19

Answer:

metals

Card 24

Chapter 19

Answer:

No, if pairs - 2 per side
for a maximum number of 8.

Card 23

Chapter 19

Answer:

Number of valence electrons

Card 25

Chapter 19

Question:

Elements on the right side of the periodic table

Card 26

Chapter 19

Question:

What charge will calcium have if it loses two electrons?

Card 27

Chapter 19

Question:

Write ion notation for Chlorine with 18 electrons

Card 28

Chapter 19

Question:

A positive ion. Occurs when a metal loses electrons (more protons than electrons)

Card 29

Chapter 19

Question:

How many electrons does Mg^{2+} have?

Card 30

Chapter 19

Question:

Write ion notation for Beryllium with 2 electrons

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Chapter 19

Question:

blank question

Card 32

Chapter 19

Question:

blank question

Card 26

Chapter 19

Answer:

2+ (losers of electrons become positive)

Card 25

Chapter 19

Answer:

non-metals

Card 28

Chapter 19

Answer:

cation

Card 27

Chapter 19

Answer:

Cl^{-}

Card 30

Chapter 19

Answer:

Be^{2+}
(4 p - 2 e = 2+ charge)

Card 29

Chapter 19

Answer:

10 electrons
(12 protons - 10 electrons = 2+ charge)

Card 32

Chapter 19

Answer:

blank answer

Card 31

Chapter 19

Answer:

blank answer